

Blackline Master 5.8

Ionic Compounds: Names and Formulas Worksheet

1. Write the formulas for the following compounds.

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|-----------------------|-------|------------------------|-------|
| (a) magnesium oxide | _____ | (k) copper(I) bromide | _____ |
| (b) sodium fluoride | _____ | (l) tin(II) iodide | _____ |
| (c) aluminum nitride | _____ | (m) iron(III) chloride | _____ |
| (d) potassium sulfide | _____ | (n) calcium phosphide | _____ |
| (e) lithium iodide | _____ | (o) lead(II) oxide | _____ |
| (f) calcium bromide | _____ | (p) lead(IV) fluoride | _____ |
| (g) beryllium oxide | _____ | (q) tin(IV) bromide | _____ |
| (h) nickel chloride | _____ | (r) copper(II) sulfide | _____ |
| (i) magnesium nitride | _____ | (s) iron(II) oxide | _____ |
| (j) aluminum sulfide | _____ | (t) calcium nitride | _____ |

2. Write the names for the following compounds.

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|-----------------------------|-------|-----------------------------|-------|
| (a) Li_2O | _____ | (k) PbS | _____ |
| (b) AlCl_3 | _____ | (l) SnO_2 | _____ |
| (c) MgS | _____ | (m) Na_2S | _____ |
| (d) CaO | _____ | (n) Mg_3P_2 | _____ |
| (e) KBr | _____ | (o) NiO | _____ |
| (f) BeF | _____ | (p) CuI | _____ |
| (g) Na_3N | _____ | (q) PbCl_4 | _____ |
| (h) Al_2O_3 | _____ | (r) FeP | _____ |
| (i) CuCl_2 | _____ | (s) CaF_2 | _____ |
| (j) FeBr_3 | _____ | (t) K_3P | _____ |